

## Compound Naming Worksheet

**Part A: Name the following binary compounds. If there is a transition element, write both names (IUPAC and the Latin).**

NaCl \_\_\_\_\_

K<sub>2</sub>S \_\_\_\_\_

CaO \_\_\_\_\_

AlBr<sub>3</sub> \_\_\_\_\_

FeI<sub>2</sub> \_\_\_\_\_

Fe<sub>2</sub>O<sub>3</sub> \_\_\_\_\_

PCl<sub>3</sub> \_\_\_\_\_

PCl<sub>5</sub> \_\_\_\_\_

NO<sub>2</sub> \_\_\_\_\_

CuS \_\_\_\_\_

Cu<sub>2</sub>S \_\_\_\_\_

S<sub>2</sub>Cl<sub>2</sub> \_\_\_\_\_

TiCl<sub>4</sub> \_\_\_\_\_

Part B: Write the formulas for the following binary compounds.

calcium oxide \_\_\_\_\_

Magnesium nitride \_\_\_\_\_

sodium selenide \_\_\_\_\_

Lithium bromide \_\_\_\_\_

Cesium chloride \_\_\_\_\_

Barium sulfide \_\_\_\_\_

Dinitrogen tetraoxide \_\_\_\_\_

Tin (II) iodide \_\_\_\_\_

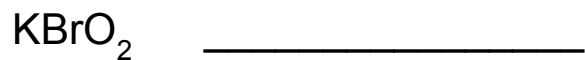
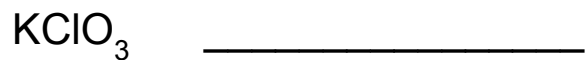
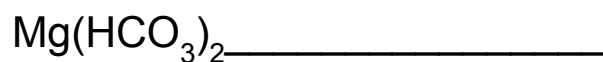
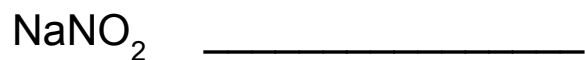
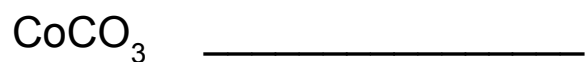
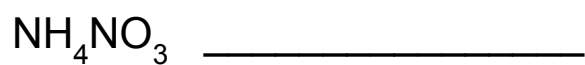
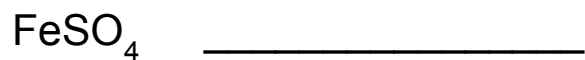
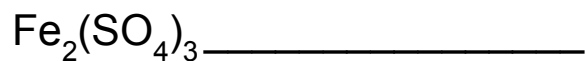
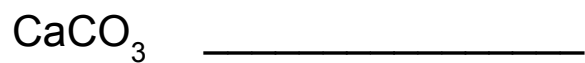
Tin (IV) oxide \_\_\_\_\_

Sulfur hexafluoride \_\_\_\_\_

Iron (III) oxide \_\_\_\_\_

Mercury (II) sulfide \_\_\_\_\_

**Part C: Name the following. If there is a transition element, write both formulas (Stock System and suffix system).**



**Part D: Write the formulas for the following.**

Tin (IV) sulfate \_\_\_\_\_

Iron (III) nitrate \_\_\_\_\_

Strontium nitrite \_\_\_\_\_

Potassium sulfite \_\_\_\_\_

Barium phosphate \_\_\_\_\_

Sodium permanganate \_\_\_\_\_

Potassium chromate \_\_\_\_\_

Sodium hydroxide \_\_\_\_\_

Monohydrogen phosphate ion \_\_\_\_\_

Magnesium dihydrogen phosphate \_\_\_\_\_

Cobalt (II) chloride hexahydrate \_\_\_\_\_

## Practice Worksheet

1. Write the names for each of the following:

a. NaBr \_\_\_\_\_

k. KNO<sub>2</sub> \_\_\_\_\_

b. Ca(OH)<sub>2</sub> \_\_\_\_\_

l. Al<sub>2</sub>S<sub>3</sub> \_\_\_\_\_

c. HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>(aq) \_\_\_\_\_

m. N<sub>2</sub>O<sub>3</sub> \_\_\_\_\_

d. NaHCO<sub>3</sub> \_\_\_\_\_

n. FeCO<sub>3</sub> \_\_\_\_\_

e. BaI<sub>2</sub> \_\_\_\_\_

o. Si<sub>2</sub>Cl<sub>6</sub> \_\_\_\_\_

f. N<sub>2</sub>S<sub>5</sub> \_\_\_\_\_

p. Cu(NO<sub>3</sub>)<sub>2</sub> \_\_\_\_\_

g. TeF<sub>4</sub> \_\_\_\_\_

q. Li<sub>2</sub>CO<sub>3</sub> \_\_\_\_\_

h. KMnO<sub>4</sub> \_\_\_\_\_

r. NaCN \_\_\_\_\_

i. (NH<sub>4</sub>)<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> \_\_\_\_\_

s. Pb(SO<sub>4</sub>)<sub>2</sub> \_\_\_\_\_

2. Write the formulas for each of the following:

a. triphosphorus nonaoxide \_\_\_\_\_ l. potassium iodide \_\_\_\_\_

b. carbon disulfide \_\_\_\_\_ m. nickel (II) chloride \_\_\_\_\_

c. potassium dichromate \_\_\_\_\_ n. aluminum hydroxide \_\_\_\_\_

d. sodium acetate \_\_\_\_\_ o. sulfur trioxide \_\_\_\_\_

e. calcium bicarbonate \_\_\_\_\_ p. iron(III) sulfite \_\_\_\_\_

f. ammonium oxide \_\_\_\_\_ q. barium chromate \_\_\_\_\_

g. copper (II) sulfate \_\_\_\_\_ r. titanium(IV) phosphide \_\_\_\_\_

h. mercury (I) nitrate \_\_\_\_\_ s. oxygen difluoride \_\_\_\_\_

i. dibromine hexaoxide \_\_\_\_\_ t. gold(III) perchlorate \_\_\_\_\_

j. diphosphorus pentaoxide \_\_\_\_\_ u.

k. potassium chlorate \_\_\_\_\_ v.

## Naming Acids

*Write the formulas for each of the following acids:*

HCl(g) \_\_\_\_\_

HNO<sub>3</sub>(aq) \_\_\_\_\_

HClO<sub>4</sub> (aq) \_\_\_\_\_

H<sub>2</sub>S (aq) \_\_\_\_\_

HBrO<sub>2</sub> (aq) \_\_\_\_\_

HPO<sub>3</sub> (aq) \_\_\_\_\_

HClO<sub>3</sub> (aq) \_\_\_\_\_

Write the names for each of the following acids:

sulfuric acid \_\_\_\_\_

carbonic acid \_\_\_\_\_

Hydrochloric acid \_\_\_\_\_

Nitric acid \_\_\_\_\_

Acetic acid \_\_\_\_\_

Hypoiodous acid \_\_\_\_\_

Bromic Acid \_\_\_\_\_