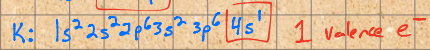
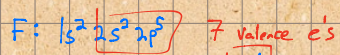


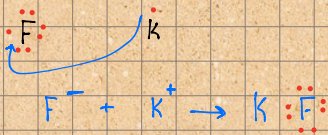
Warm-up:

1. Write the electron configurations for Fluorine and Potassium
2. Determine the # of valence e^-

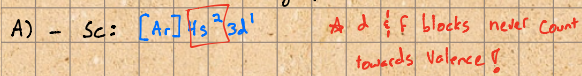


- Octet Rule: all atoms will either gain or lose electrons to obtain a full valence shell [8 valence e^-]
↳ Two exceptions → H & He hold 2 valence e^-

• Electron Dot Diagrams [Lewis Dot Diagram]: illustrates valence e^-



Ex 2: Write the Noble Gas Config for Scandium:



B) Determine the most likely common charge of Sc. Explain!

