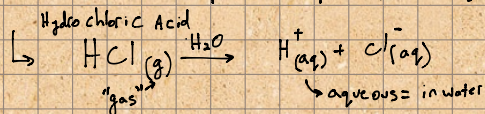


## Warm-up:

- Calcium and Chlorine form an ionic compound.
  - Draw a Lewis structure (dot diagram) of the most likely compound to form
  - What is the ionic formula of this compound?
  - What is the name of this compound?
- Write the chemical formula for each compound:
  - potassium oxide
  - copper (II) sulfide
- Write the chemical name for the following compounds:
  - $\text{CoI}_2$
  - $\text{BaO}$

## • Acids & Bases

- Acid: molecular compound that produces  $\text{H}^+$  ions in a solution [in  $\text{H}_2\text{O}$ ]



### Naming of Acids:

- If the anion ends in "ide" then **hydro-STEM-ic**



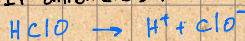
Hydro-Iod-ic = Hydroiodic acid

- If anion ends in "ate" then **STEM-ic Acid**

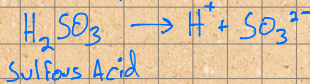


Nitric Acid Nitrate

- If anion ends in "ite" then **STEM-ous Acid**



Hypochlorous Acid hypochlorite



### 6 Most Common Acids:

- $\text{HCl}$  = hydrochloric acid
- $\text{HNO}_3$  = Nitric Acid
- $\text{H}_2\text{SO}_4$  = Sulfuric Acid
- $\text{H}_2\text{CO}_3$  = Carbonic Acid
- $\text{H}_3\text{PO}_4$  = Phosphoric Acid
- $\text{H}_2\text{C}_2\text{H}_3\text{O}_2$  = Acetic Acid

- **Bases:** produce hydroxide [ $\text{OH}^-$ ] when dissolved in  $\text{H}_2\text{O}$

↳ Most are ionic compounds → Follow these rules



Sodium Hydroxide